

(LM 2007)

MARCH 2018

Sub. Code: 2007

B.PHARM. DEGREE EXAMINATION
PCI Regulation SEMESTER – I
PAPER IV – PHARMACEUTICAL INORGANIC CHEMISTRY

Q.P. Code: 562007

Time: Three hours

Maximum: 75 Marks

I. Elaborate on: **(2 x 10 = 20)**

1. Give the principle reaction involved in the limit test for Arsenic with a neat diagram of the apparatus used for it.
2. Define Radioactivity. How to measure radioactivity and explain the storage condition and precaution to be followed when handling radioactive substance?
3. Give the functions of major physiological ions used as electrolyte in the replacement therapy. Give the composition and uses of Oral rehydration salt.

II. Write notes on: **(7 x 5 = 35)**

1. What are antidotes? Give the preparation, properties, assay and uses of Sodium thiosulphate.
2. What are buffers? Give the types of buffers, preparation and stability of buffers used in pharmaceutical substances.
3. What are properties of α , β and γ rays?
4. What are Anti-microbials? Write the preparation and assay of Hydrogen peroxide.
5. Brief account about the Iodine and its solution.
6. What are expectorants? Give the preparation, properties, assay and uses of ammonium chloride.
7. Brief history of Indian Pharmacopeia.
8. Write the sources of impurities in pharmaceutical substances.
9. Write about the principle and reaction involved in the limit test of chlorides.

III. Short answers on: **(10 x 2 = 20)**

1. Define radio isotopes.
2. Write about the formula of any two emetics.
3. Define cathartics and give the formula of sodium orthophosphate.
4. Write about the formula of any two antacids.
5. Write about the formula, properties and uses of Ferrous Gluconate.
6. What is the use of activated Charcoal?
7. What are dentrifices? Give the role of fluorides in dental products.
8. Define isotonicity.
9. Write about the formula, properties and uses of Calcium gluconate.
10. What are acidifiers with two examples?

(LN 2007)

SEPTEMBER 2018

Sub. Code: 2007

B.PHARM. DEGREE EXAMINATION
PCI Regulation SEMESTER – I
FIRST YEAR
PAPER IV – PHARMACEUTICAL INORGANIC CHEMISTRY

Q.P. Code: 562007

Time: Three hours

Maximum: 75 Marks

I. Elaborate on: Answer any TWO questions. (2 x 10 = 20)

1. a) Write the principle and reaction involved in the limit test for Iron.
b) Write a note on Indian Pharmacopoeia.
2. a) Write about role of fluoride in the treatment of dental caries.
b) Write a note on Zinc eugenol cement.
3. a) Define and classify antimicrobial with example.
b) Write about the preparation, assay and uses of chlorinated lime.

II. Write notes on: Answer any SEVEN questions. (7 x 5 = 35)

1. Explain the methods of adjusting isotonicity.
2. Write about the preparation, assay and uses of calcium gluconate.
3. Explain the principle and reaction involved in the limit test for lead.
4. Define alum and gives the formula, properties and uses of potash alum.
5. Write a note on Sodium iodide - ^{131}I .
6. Write about the preparation, assay and uses of copper sulphate.
7. Define antacid and gives a brief note on combination of antacids.
8. Briefly describe physiological acid base balance.
9. Discuss about measurement of radio activity.

III. Short answers on: Answer ALL questions. (10 x 2 = 20)

1. Define Haematinics and give example.
2. Write the formula and uses of potassium chloride.
3. Define emetics and give one example.
4. Define Half life period.
5. Principle of limit test for sulphate.
6. Write the principle involved in assay of ammonium chloride.
7. Define astringent and give example.
8. Write about Poison.
9. Define buffer capacity.
10. Write the composition of oral rehydration salt?

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MARCH 2019

Sub. Code: 2007

B.PHARM. DEGREE EXAMINATION
PCI Regulation SEMESTER – I
PAPER IV – PHARMACEUTICAL INORGANIC CHEMISTRY

Q.P. Code: 562007

Time: Three hours

Maximum: 75 Marks

I. Elaborate on: Answer any TWO questions. (2 x 10 = 20)

1. Describe in detail about buffered isotonic solutions, methods of adjusting tonicity and measurement of tonicity.
2. a) Write a brief note on sources of impurities in pharmaceuticals.
b) Write the principle and reaction involved in the limit test for lead.
3. Define cathartic and discuss about the properties and uses of the followings:
a) Magnesium sulphate b) Sodium ortho phosphate c) Kaolin d) Bentonite

II. Write notes on: Answer any SEVEN questions. (7 x 5 = 35)

1. Discuss about the storage condition, precaution and handling of radioactive materials.
2. Write about the preparation, assay and uses of ammonium chloride.
3. Explain the principle and reaction involved in the limit test for iron.
4. Define alum and give the formula, properties and uses of calcium carbonate.
5. Write a note on emetics.
6. Write about the preparation, assay and uses of ferrous sulphate.
7. Define radioactivity and give a brief note on alpha, beta and gamma radiation.
8. Write about the preparation, assay and uses of chlorinated lime.
9. Discuss about the preparation, properties and uses of sodium bicarbonate.

III. Short answers on: Answer ALL questions. (10 x 2 = 20)

1. Write the formula and uses of ferrous sulphate.
2. Write the assay of sodium thiosulphate.
3. Define cathartic and give example.
4. Write the principle involved in assay of sodium chloride.
5. Principle of limit test for chloride.
6. Define expectorant and give one example.
7. Give the assay of copper sulphate.
8. Write the formula and uses of potassium permanganate.
9. Define pharmacopoeia.
10. Write the ideal properties of antacid.

B.PHARM. DEGREE EXAMINATION
PCI Regulation SEMESTER – I
PAPER IV – PHARMACEUTICAL INORGANIC CHEMISTRY

Q.P. Code: 562007

Time: Three hours

Maximum: 75 Marks

I. Elaborate on: Answer any TWO questions. (2 x 10 = 20)

1. Discuss in detail about the apparatus and principle involved in the limit test for Arsenic.
2. Classify Antacid and write a note on acid neutralizing capacity of Aluminium Hydroxide gel. Give the preparation, assay and properties of any one Antacid.
3. a) Explain the role of Electrolytes in acid base balance.
b) Give the preparation, assay and uses of Sodium chloride.

II. Write notes on: Answer any SEVEN questions. (7 x 5 = 35)

1. Define Antidotes and write a note on sodium nitrite.
2. What are the precautions to be followed while handling radioactive materials?
3. Write notes on combination of Antacids.
4. Write the principle involved in the limit test for Iron.
5. Write the preparation, properties, assay and uses on Hydrogen peroxide.
6. Define the terms with examples.
a) Antacid b) Astringent c) Laxative d) Antiseptic e) Disinfectant.
7. Define Haematinics. Write the method of preparation, assay and uses of Ferrous Sulphate.
8. Give the precautions and pharmaceutical applications of radioactive substances.
9. Describe about the Calcium carbonate.

III. Short answers on: Answer ALL questions. (10 x 2 = 20)

1. Define dentifrices with examples.
2. Write a note on assay of ammonium chloride.
3. Write the composition of Ringer's solution.
4. Osmotic laxative.
5. Define antidotes with examples.
6. What are official compounds of iron?
7. Radio opaque contrast medium.
8. Write the role of fluoride in the dental caries.
9. Boric acid + Glycerol →.
10. Write a note on alum.

[BPHARM 0921]

SEPTEMBER 2021
(SEPTEMBER 2020 EXAM SESSION)

Sub. Code: 2007

B.PHARM. DEGREE EXAMINATION
PCI Regulation 2017 – SEMESTER I
PAPER IV – PHARMACEUTICAL INORGANIC CHEMISTRY
Q.P. Code: 562007

Time: Three hours

Maximum: 75 Marks

I. Elaborate on: Answer any TWO questions. (2 x 10 = 20)

1. What are Limit tests? Discuss in detail about the principle involved in the limit test of Iron.
2. Discuss in detail the different methods used in the measurement of Radioactive substances with a note on their storage conditions.
3. Write a detailed note on electrolytes used in Replacement therapy and importance of Oral Rehydration Salt (ORS).

II. Write notes on: Answer any SEVEN questions. (7 x 5 = 35)

1. What are Antacids? Give the preparations, properties and uses of Sodium bicarbonate.
2. Give the principle, and reaction involved in limit test of Chloride.
3. Role of Fluorides in treatment of Dental caries.
4. What are Antidotes? Give the preparation, properties and use of Sodium thio sulphate.
5. Define the terms: (a)Expectorants (b)Laxatives (c)Astringents (d)Disinfectants.
6. Give the properties and uses of the following. (a)Kaolin (b)Sodium ortho phosphate.
7. Write short notes on Iodine and its preparations.
8. What are Emetics ? Give the preparation , properties and use of Copper sulphate.
9. Write a short note on sources of impurities in Pharmaceutical substances.

III. Short answers on: Answer ALL questions. (10 x 2 = 20)

1. What are Electrolytes?
2. Any two applications of Radioisotopes.
3. Give the properties and use of Ferrous sulphate.
4. Write a note on Alum.
5. Define Isotonicity.
6. What are Acidifiers?
7. Preparation and uses of Calcium carbonate.
8. What are Haematinics?
9. Define Buffer capacity.
10. Properties and use of Potassium permanganate.